

Comparison of field and laboratory low-pH Fe(II) oxidation rates

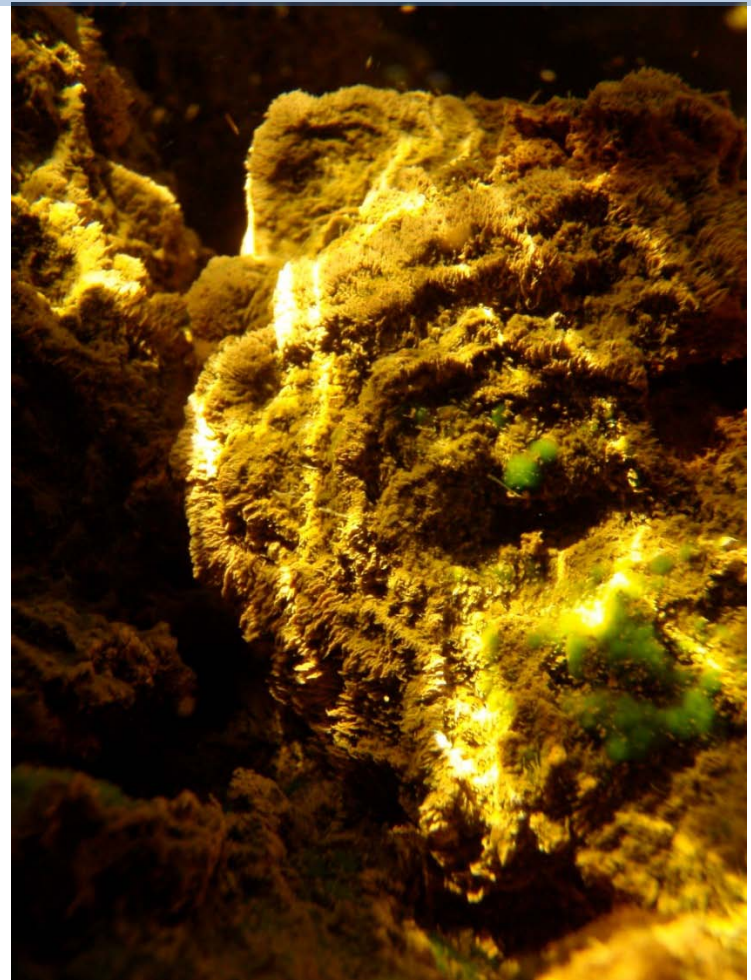
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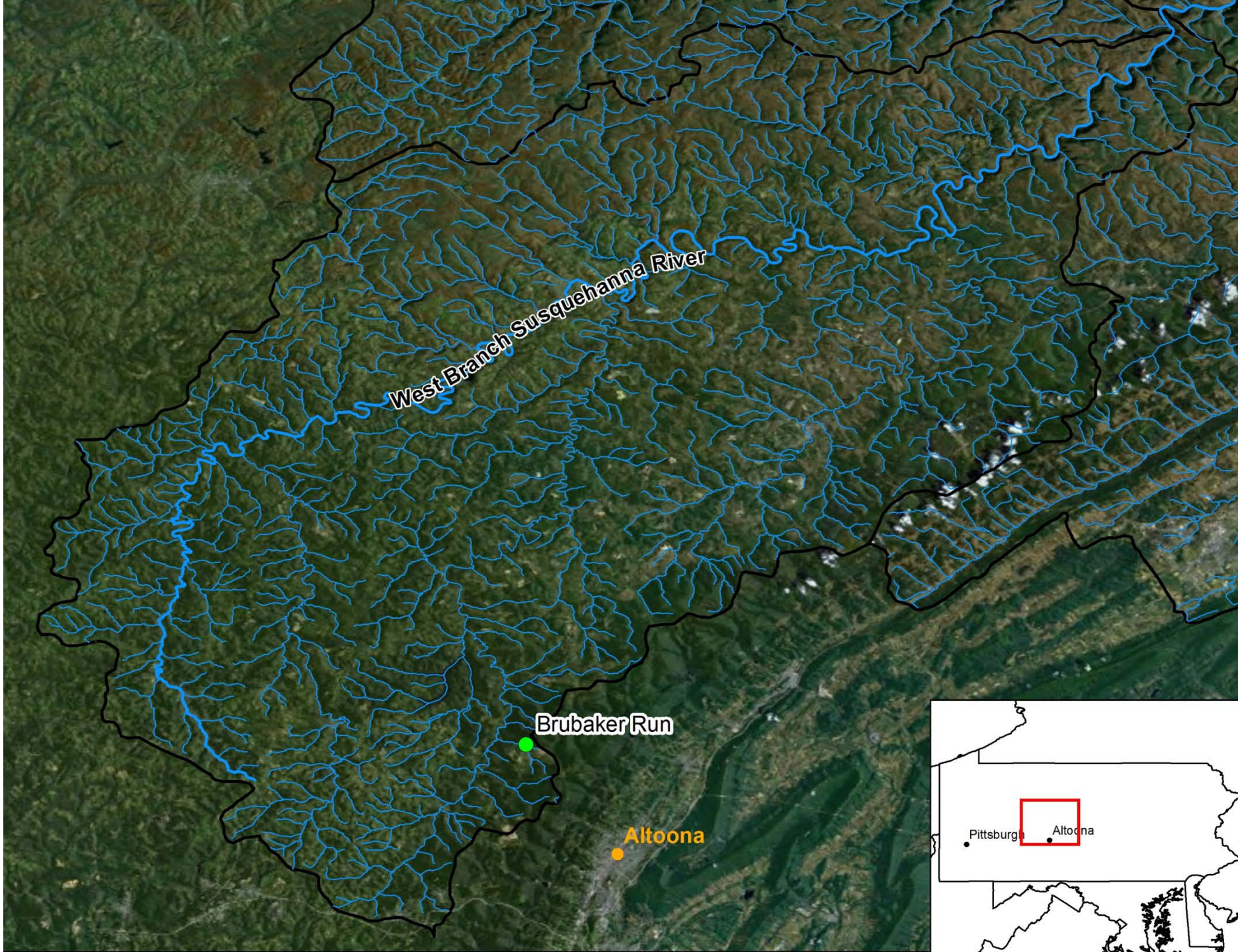




Motivation for Research

- The estimated worldwide liability costs associated with the current and future remediation of acid drainage are approximately \$100 billion (Hudson-Edwards et al. 2012).
- Estimated between 95 – 99.999% of total mined material for metal mining becomes waste (Nordstrom 2012).
- Coal mine drainage (**CMD**) is responsible for 10,000 km of streams in the Appalachian region, Northeastern USA (Herlihy et al. 1990).





Coal Mining Operations

West Branch Susquehanna River

Brubaker Run

Altoona



**Coal Mining
Operations**

**EPA Impaired
Streams**

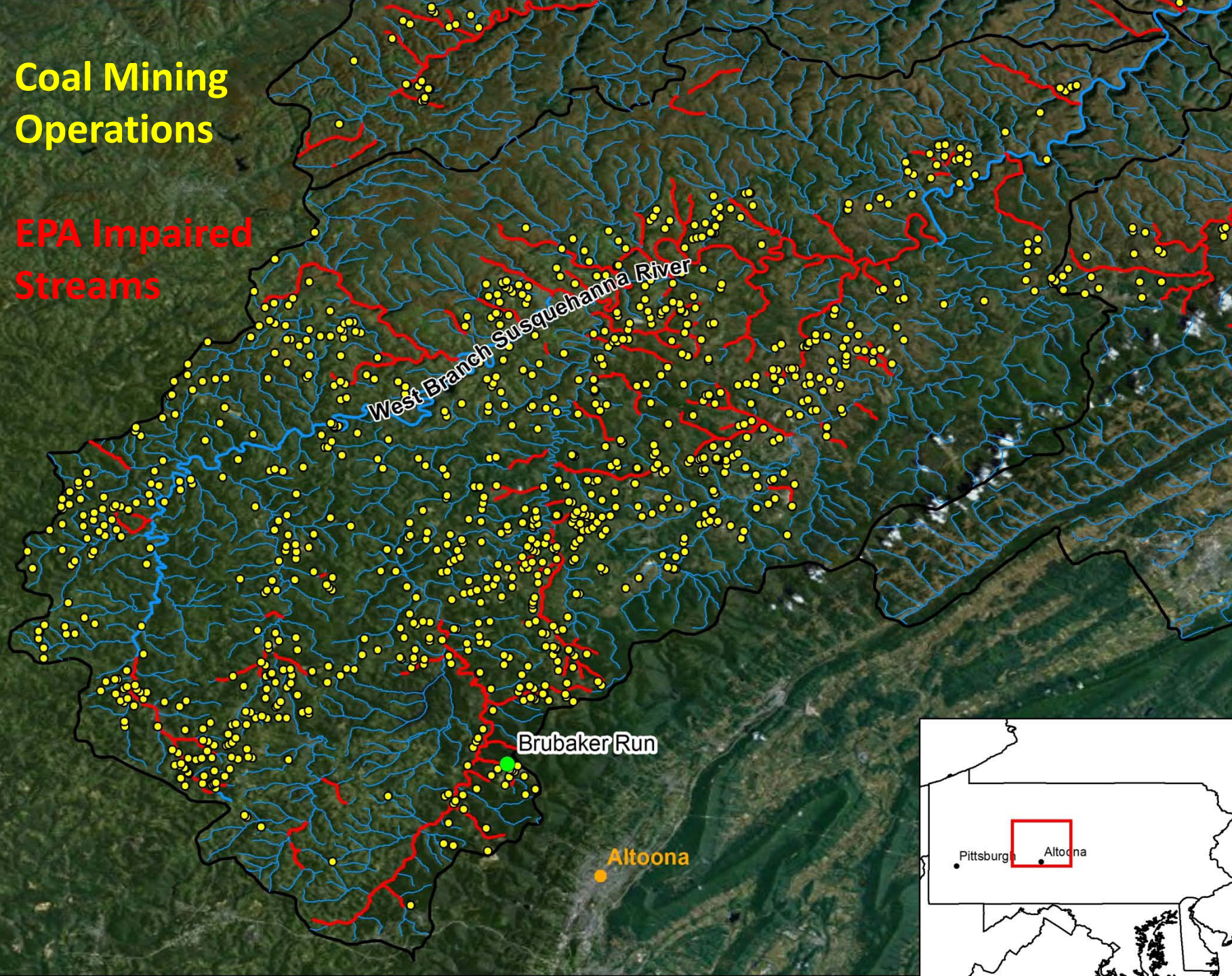
West Branch Susquehanna River

Brubaker Run

Altoona

Pittsburgh

Altoona



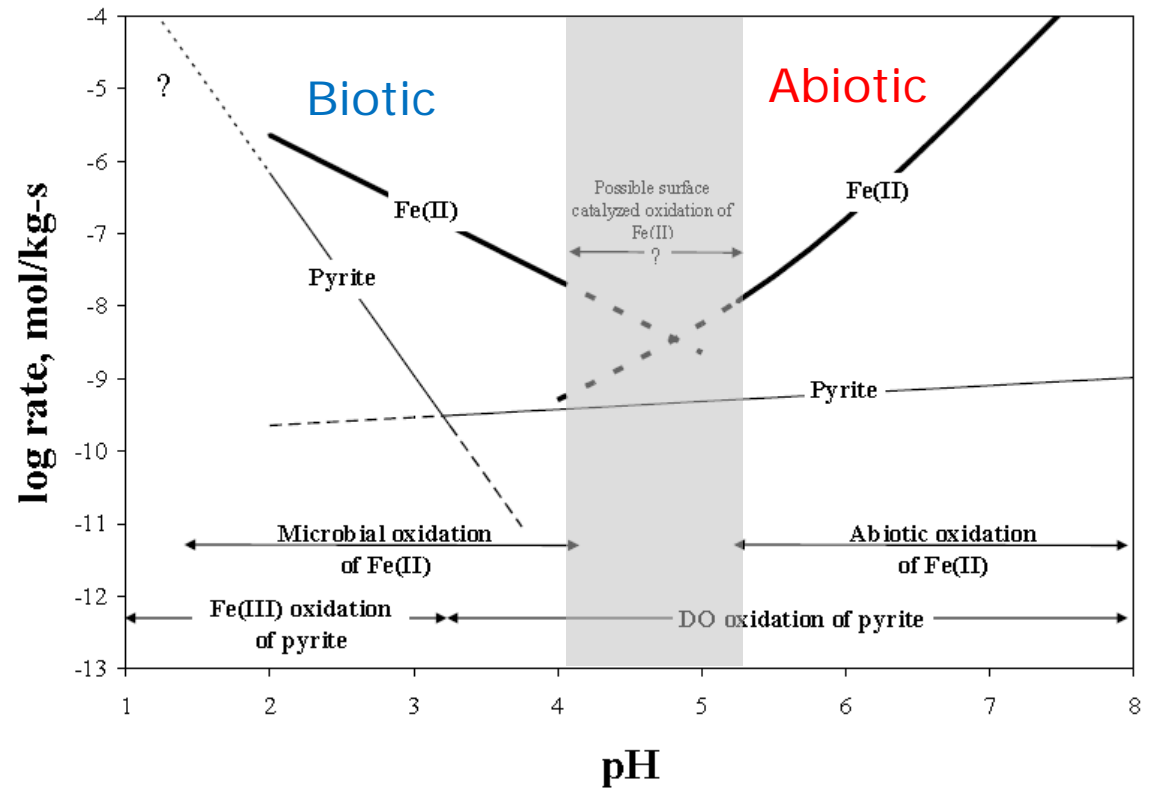
Fe(II) Oxidation Kinetics – Brief Overview

Fe(II) – **Soluble**,
dissolved in solution

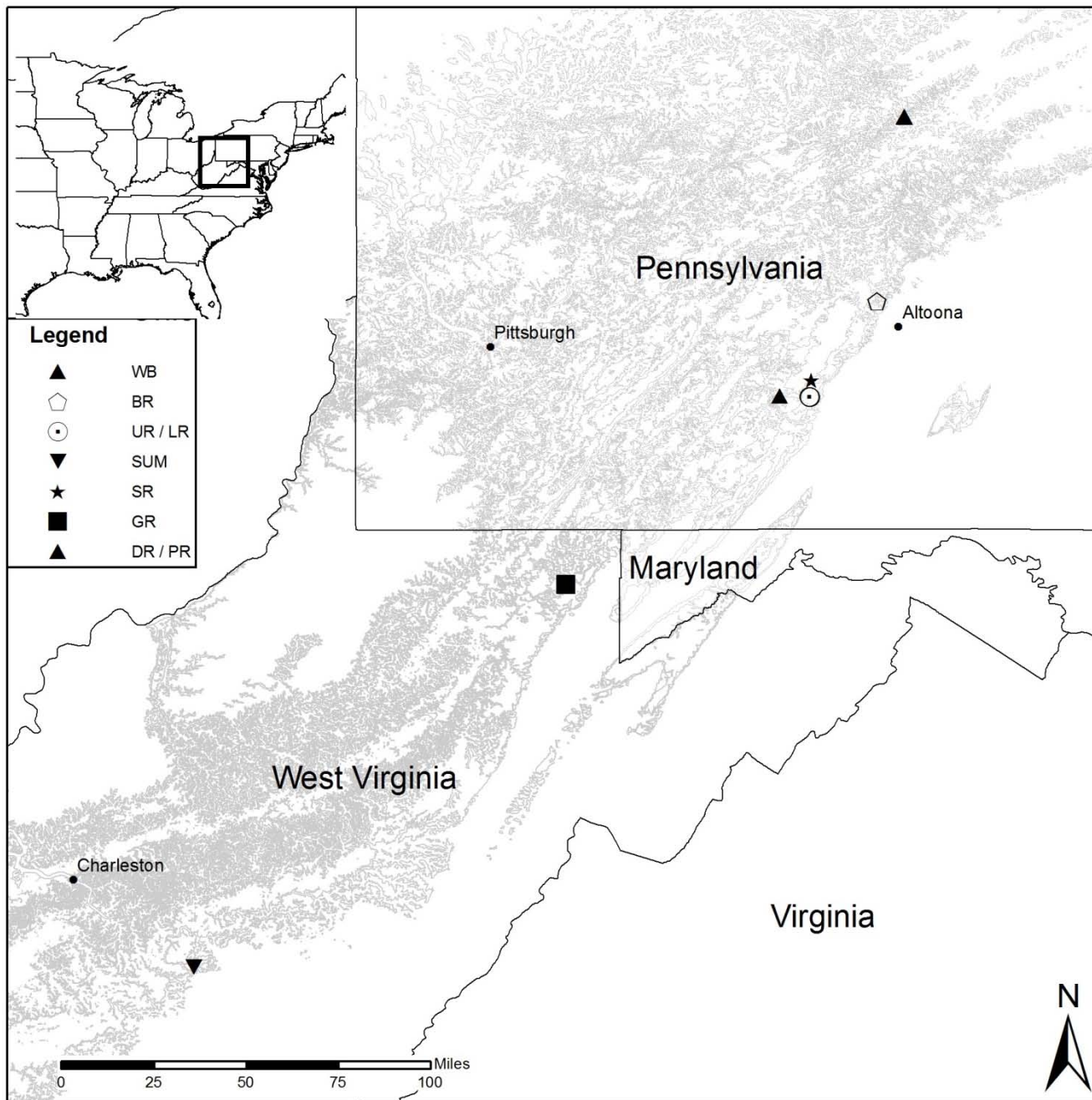
Fe(III) – **Insoluble**,
precipitates out of solution

Goal – Oxidize and
precipitate

Fe(II) → Fe(III)



Williamson et al., 2006



Fe(II) oxidation Rates?



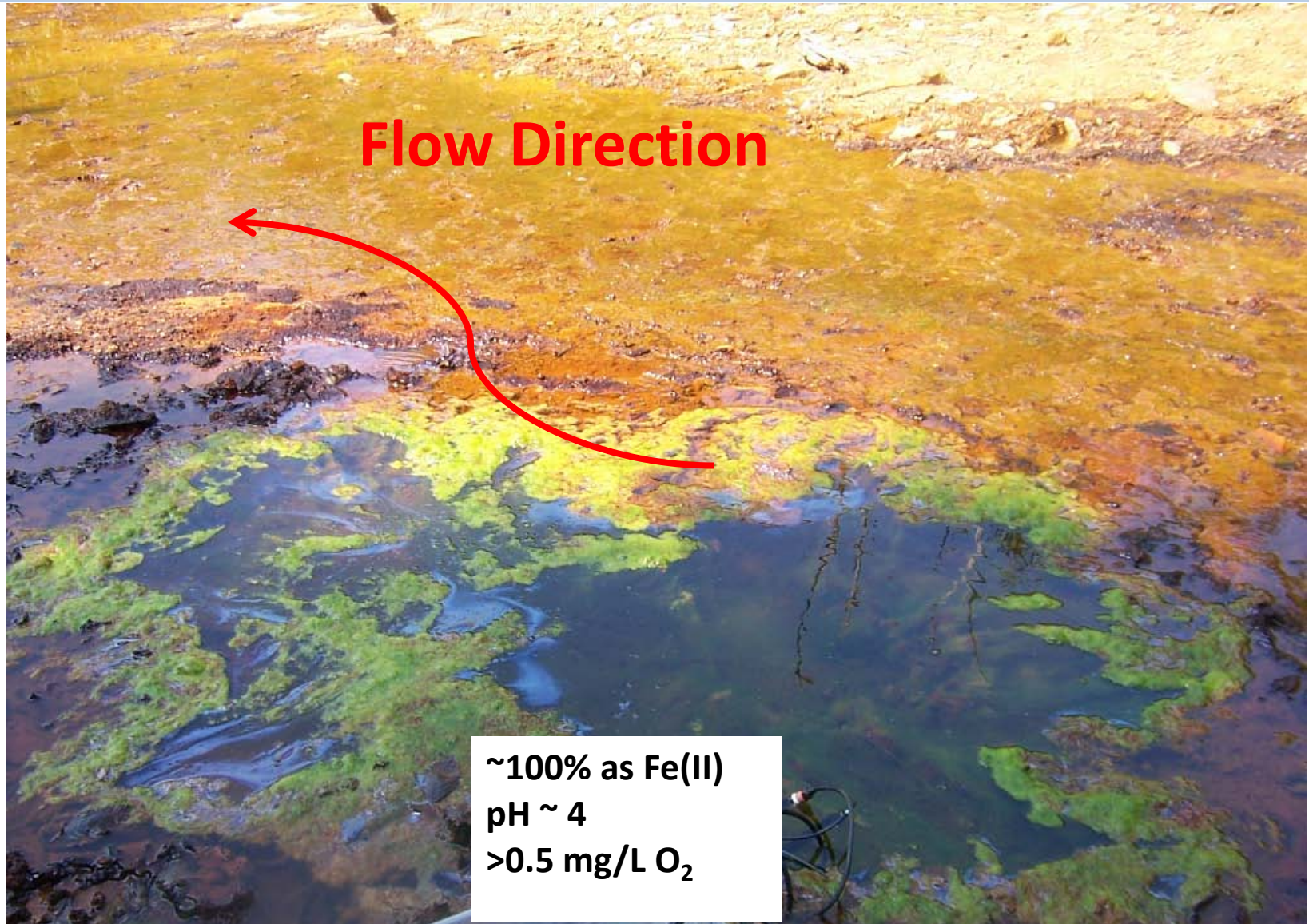
Research Approach

- Establish Fe(II) oxidation rates in the field
- Measure laboratory Fe(II) oxidation rates via constant flow reactors
- Establish specific properties of Fe sediments – ie: active microbial biomass and community structure

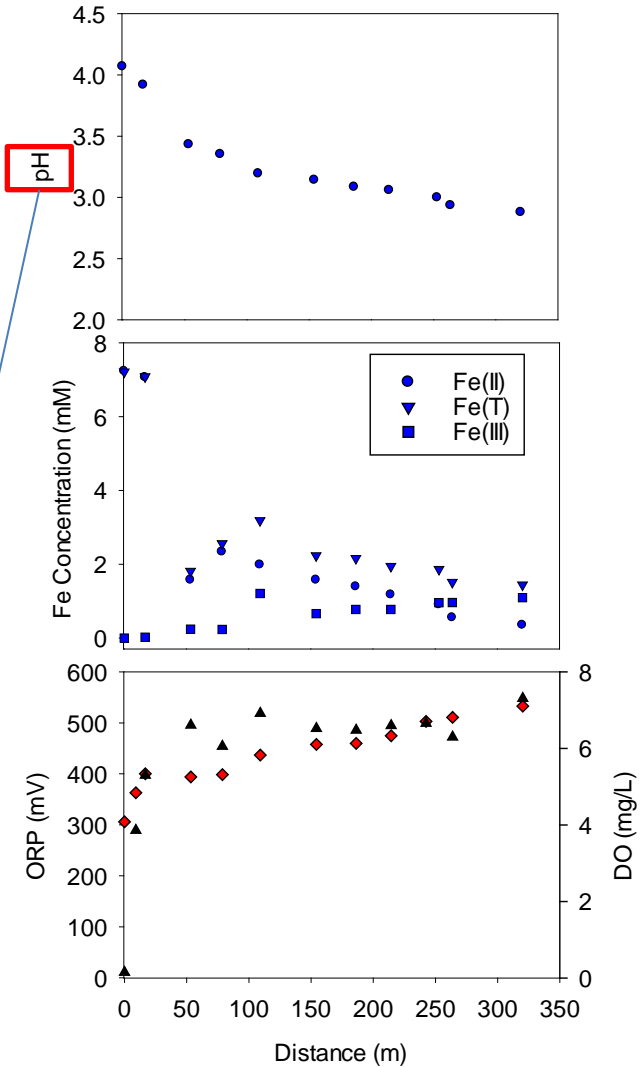
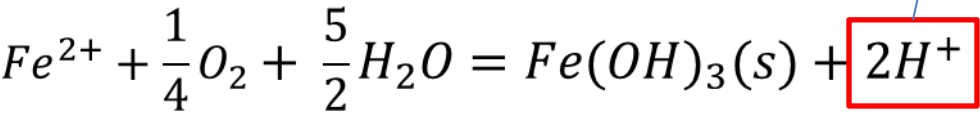
Goal:

Develop a Fe(II) oxidation design process to optimize land use for passive treatment, prior to alkaline generating processes

Emergence Source Chemistry



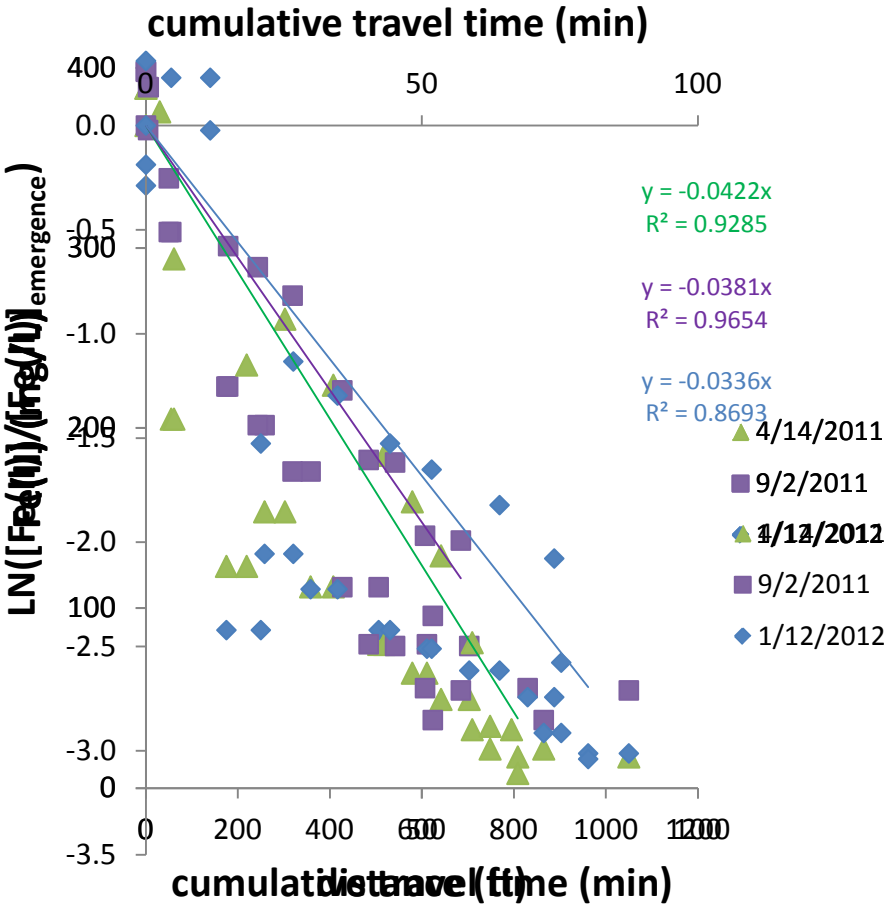
Downstream Source Chemistry



Measurement of Field Fe(II) Oxidation Rates

- Geochemical Parameters measured as a function of distance – pH, ORP, Fe(II), Fe(T), DO
- Concurrent measurements of velocity where made to establish a hydraulic residence time.

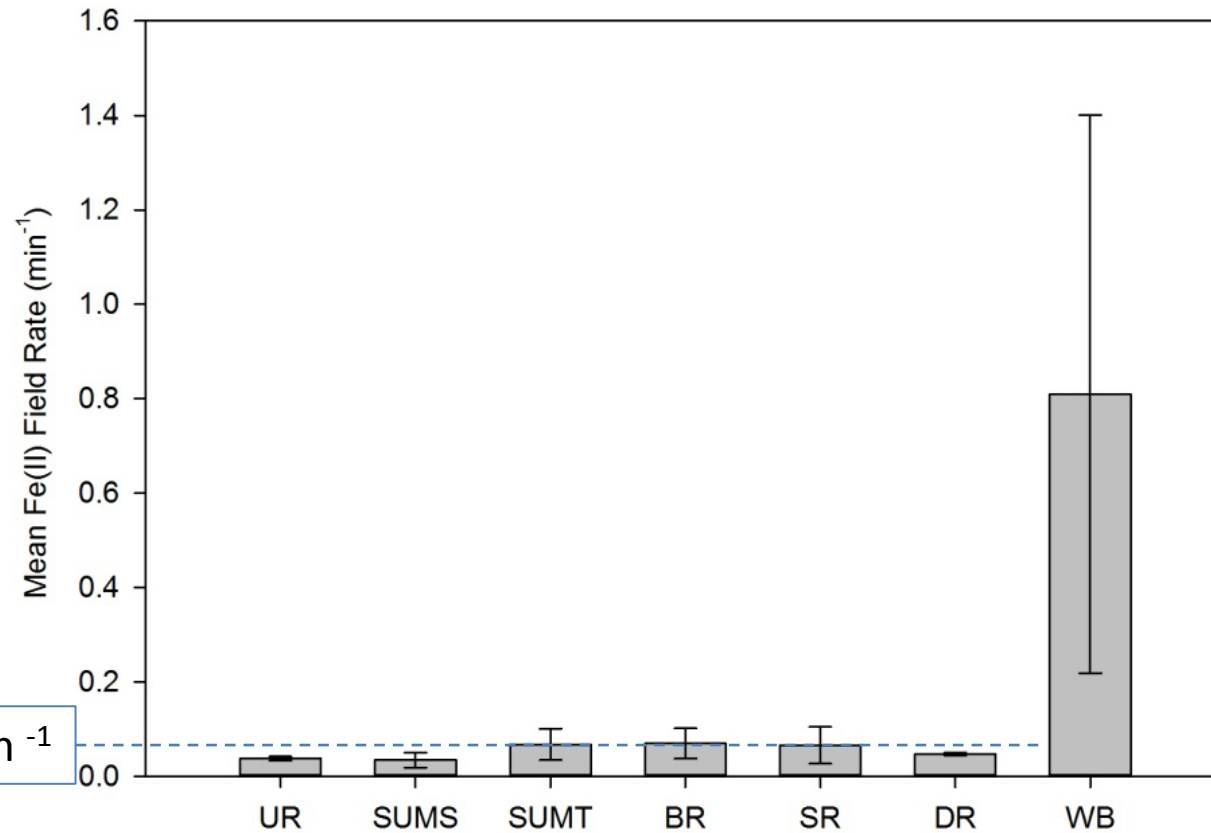
$$travel\ time\ (min) = \frac{\bar{d}}{\left(\frac{v_n + v_{n+1}}{2}\right)}$$



$$C(t) = C_o e^{-kt}$$



Field Fe(II) Oxidation Rates Summary



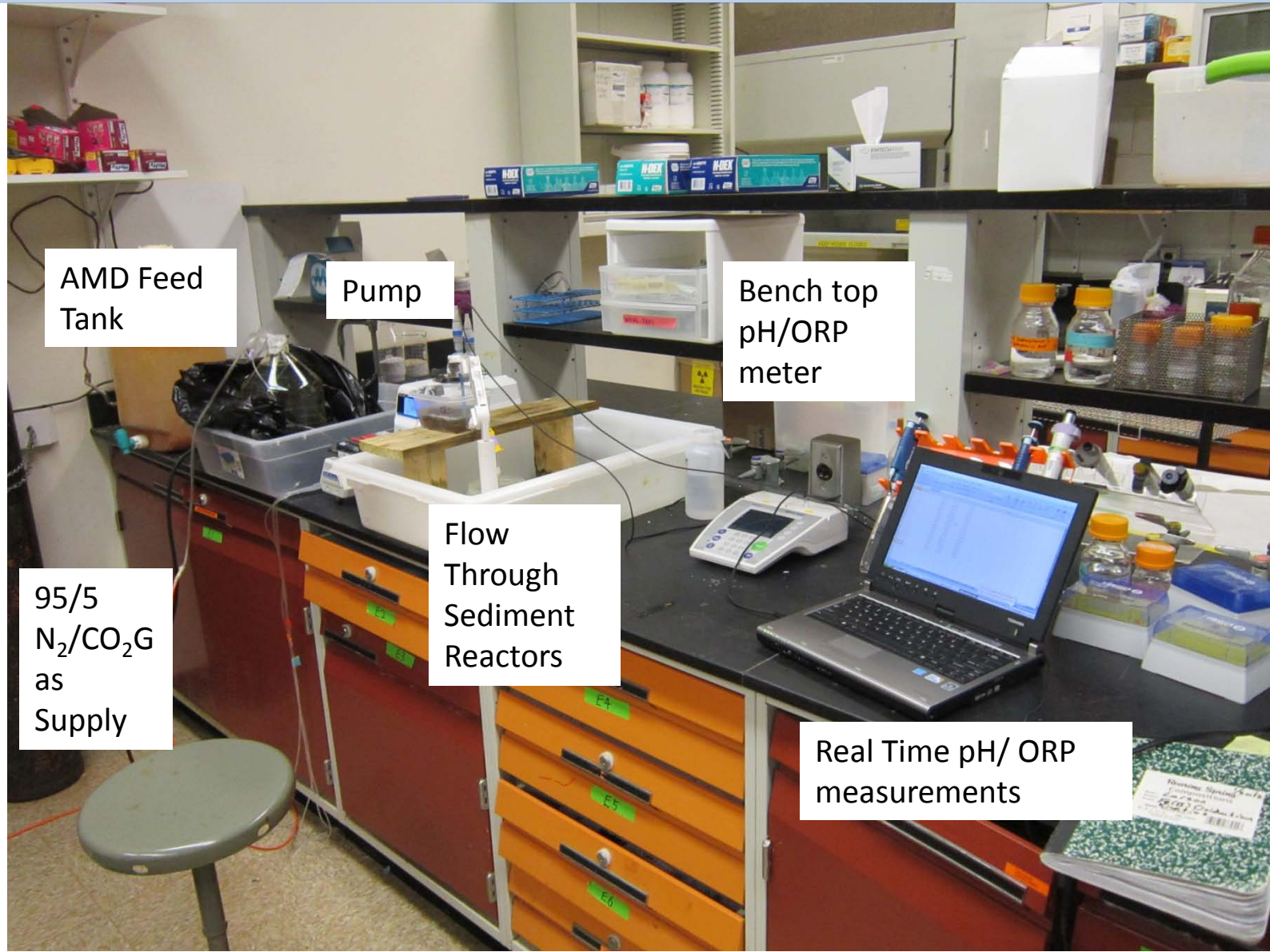
$$k_{\text{Fe(II)}} = 0.054 \text{ min}^{-1}$$

Laboratory Fe(II) Oxidation Reactors

- Sediments reactors run in triplicate (3)
- Simultaneous 'no sediment' control reactors (2)
- 8 hour residence time
- Constant flow rate confirmed gravimetrically
- Experiment duration ~60 hours

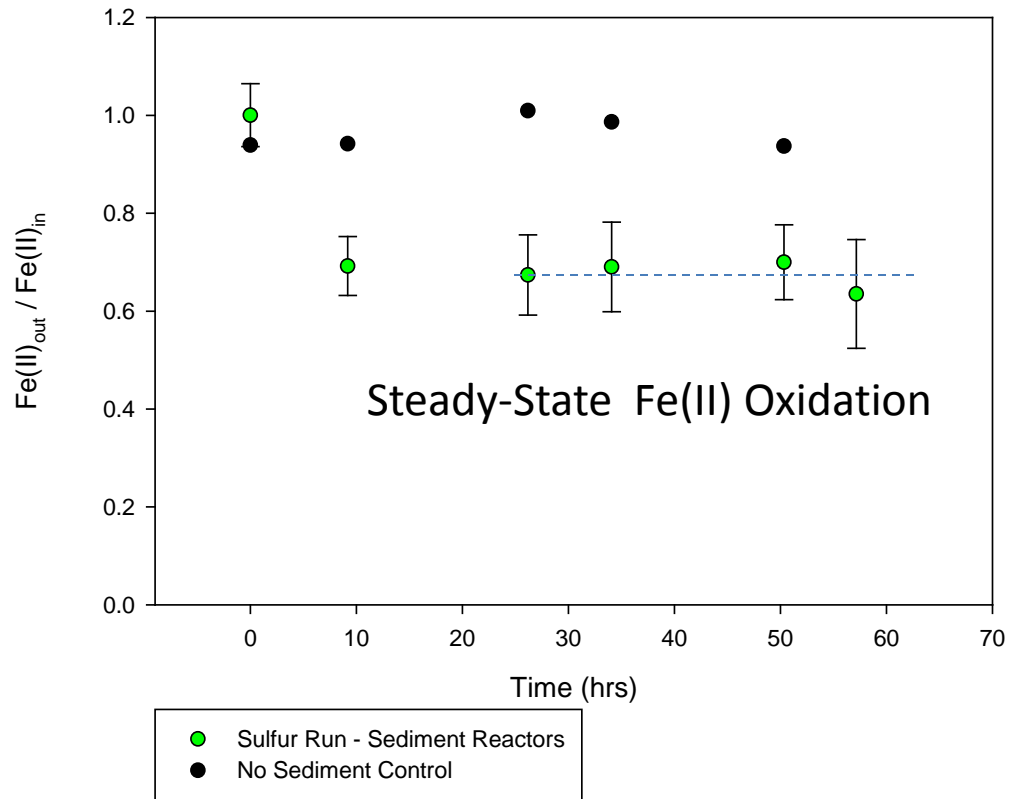


Laboratory Fe(II) Oxidation Rates : Experimental Set-up





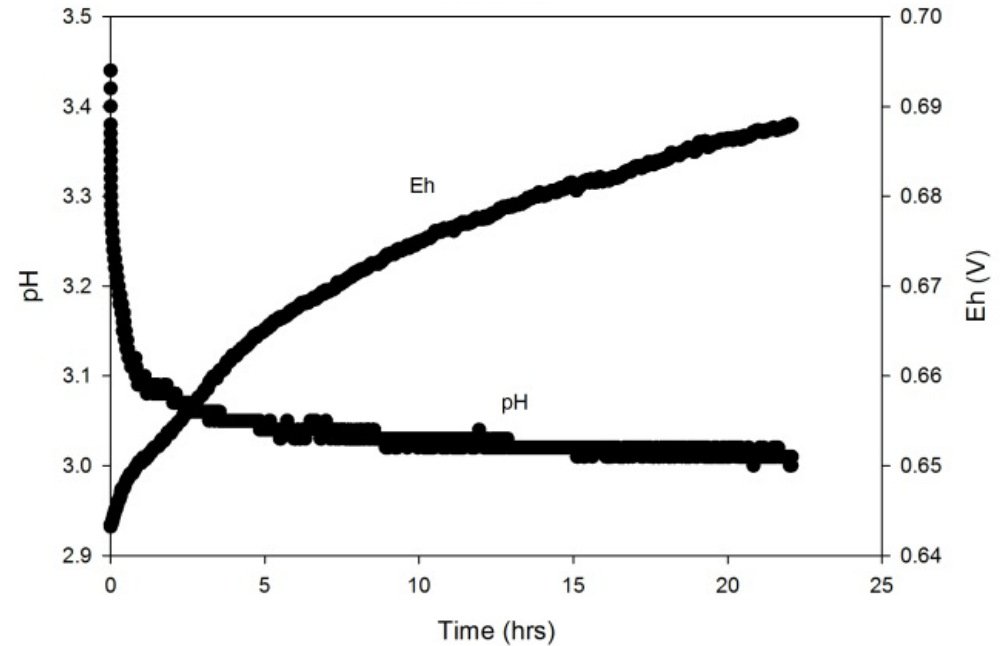
Laboratory Fe(II) Oxidation Rates



Rate constants were normalized to sediment mass and subsequent active biomass



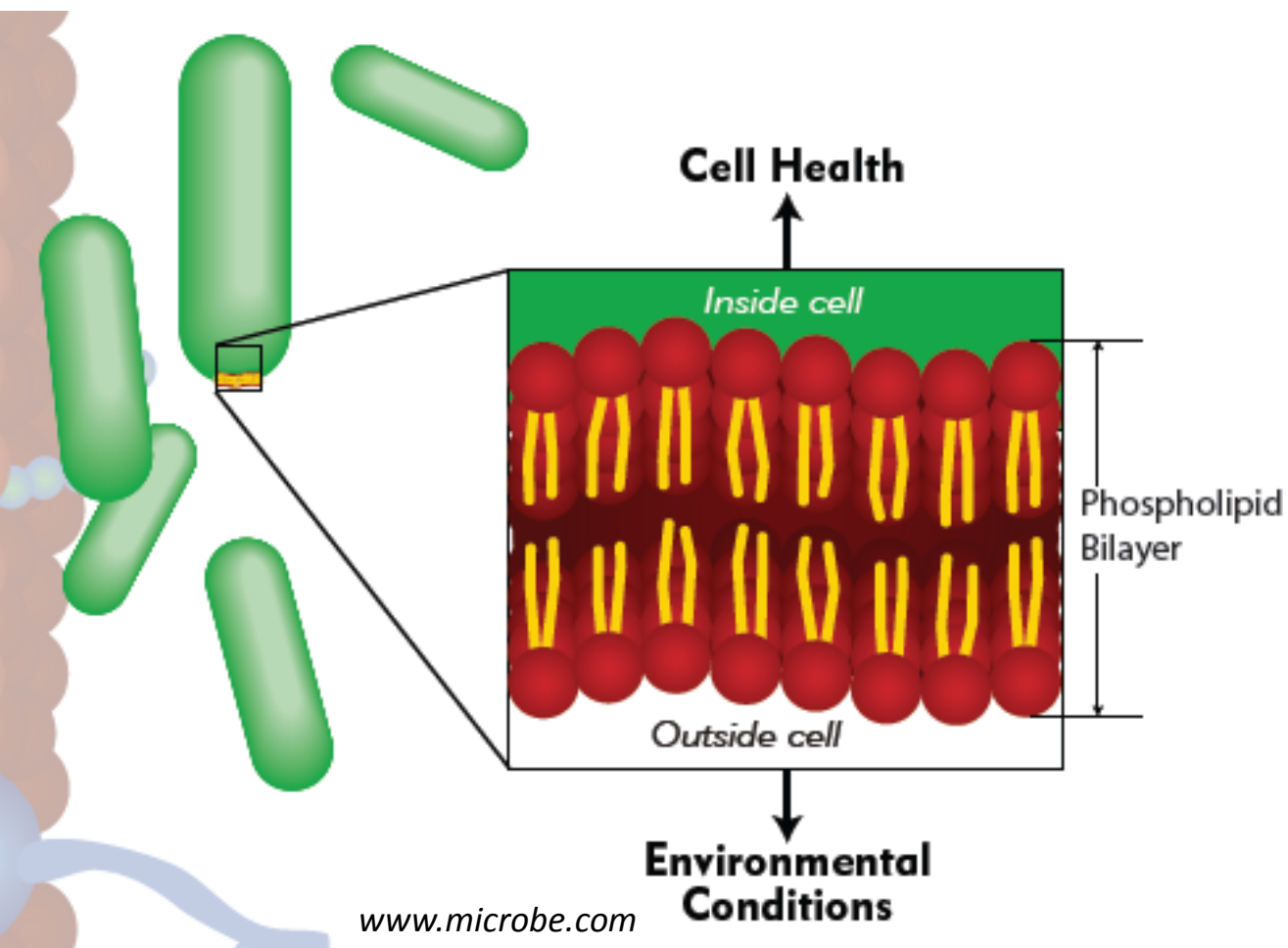
Laboratory Fe(II) Oxidation Rates - Real Time pH/ORP



Geochemical parameters

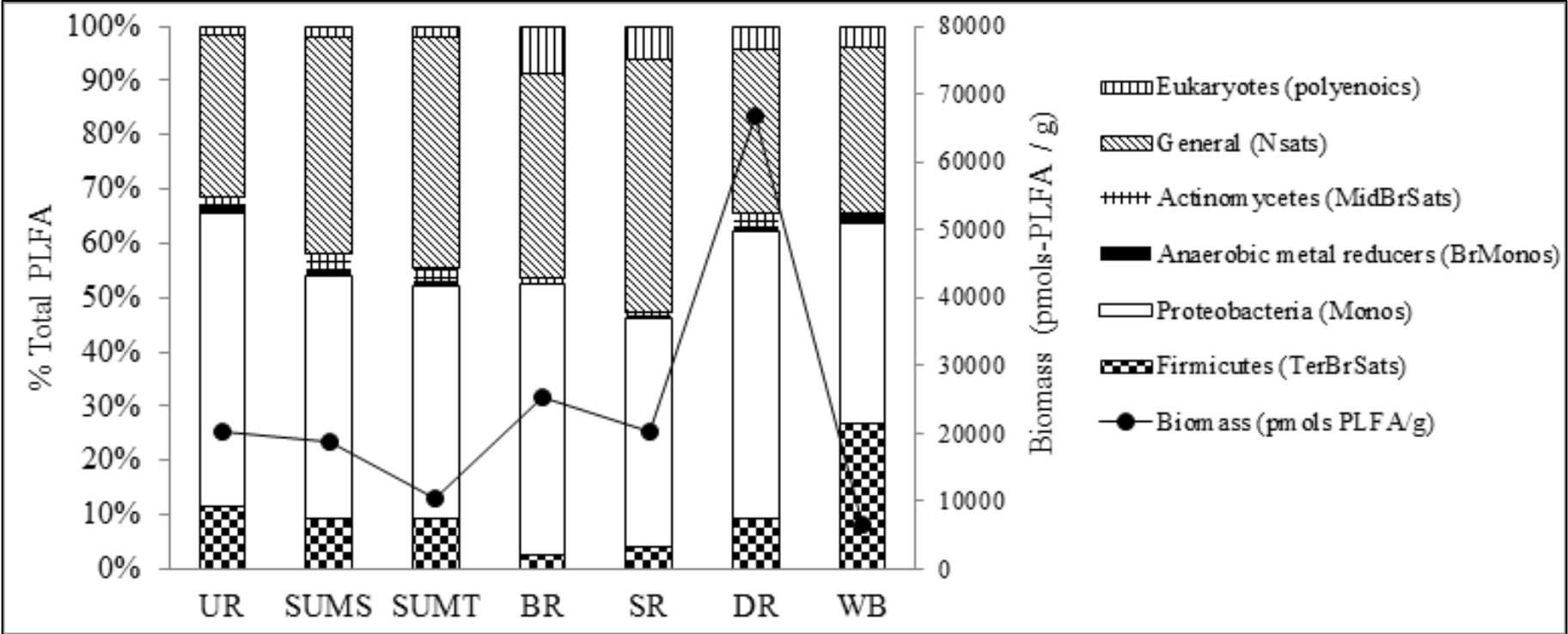
pH/ORP recorded in real-time
every 30 seconds for ~23 hours

Phospholipid Fatty Acids (PLFA)

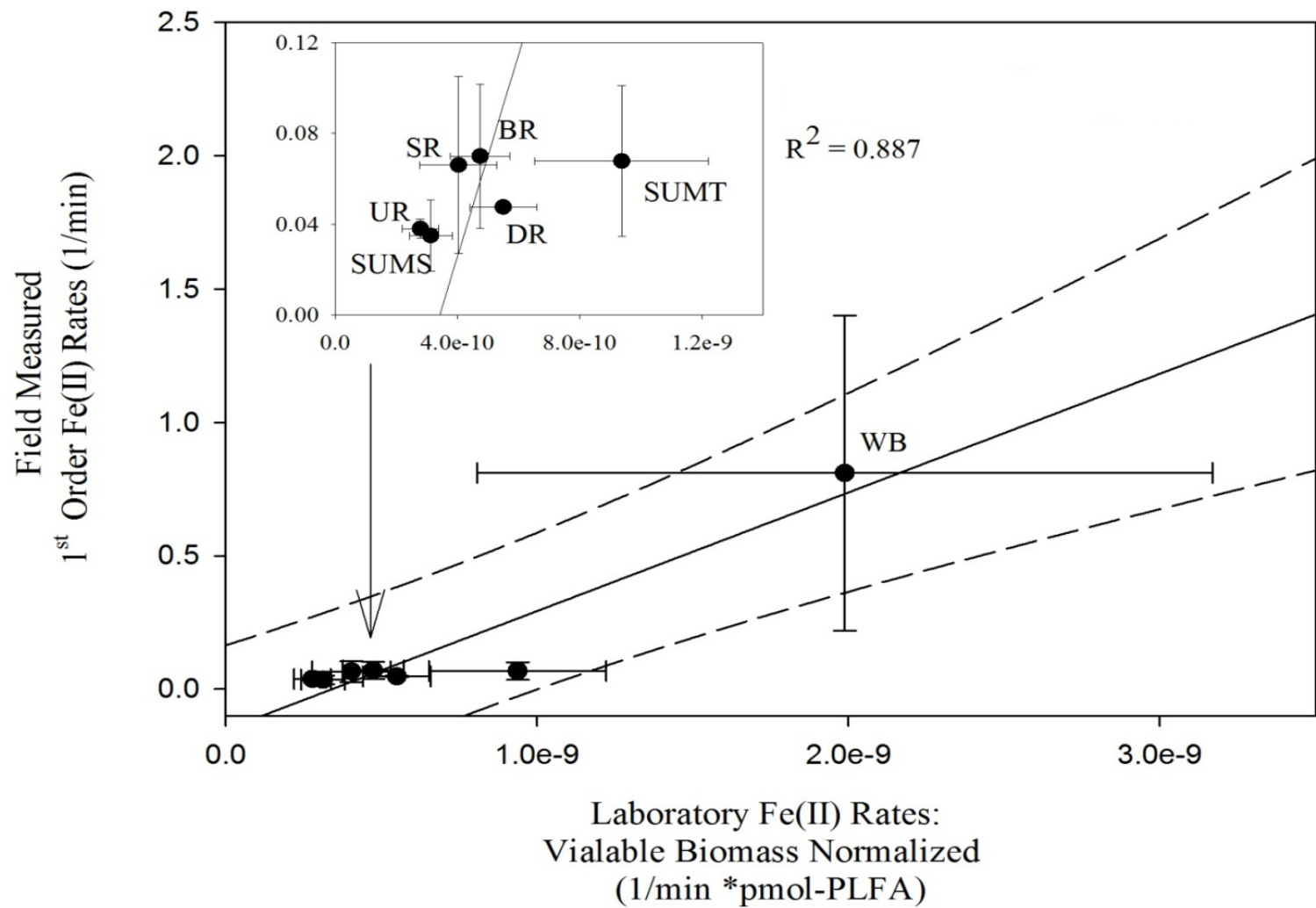


- 'Active' biomass concentration, cell membranes decay rapidly upon death
- Microbial communities have specific PLFA profiles (ie; Proteobacteria, Anaerobic metal reducers, eukaryotes)

Phospholipid Fatty Acids (PLFA)



Comparison of field and laboratory Fe(II) oxidation rates



Bioremediation Project – Summerlee, WV



*Courtesy of Levi Rose, Plateau Action Network

Bioremediation Project – Summerlee, WV



Baffles increase hydraulic residence time



*Courtesy of Levi Rose, Plateau Action Network

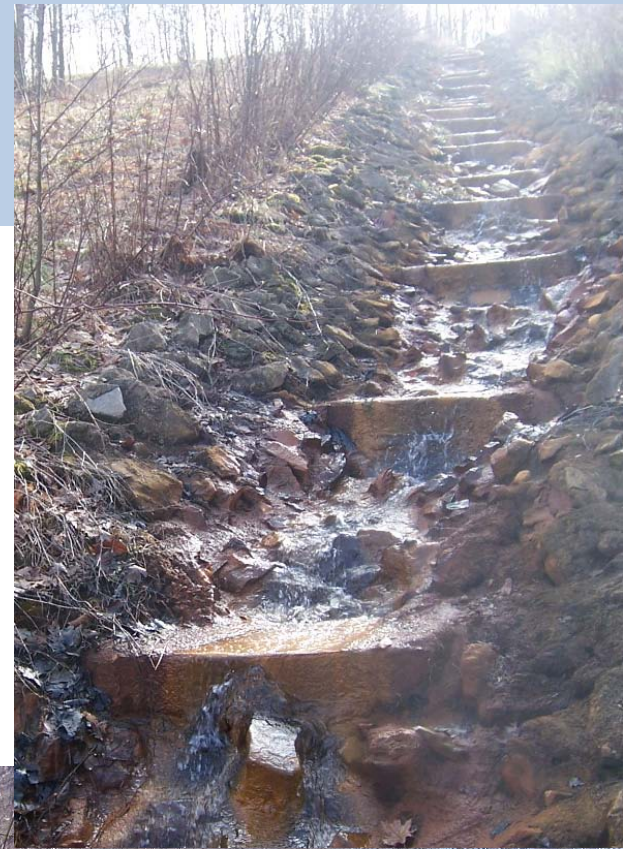
Bioremediation Project – Summerlee, WV

Pollutant	Pre-Construction Loading (lbs/yr)	Post- Construction Loading (lbs/yr)	Reduction
Acid	323,701	214,254	34%
Iron	75,385	48,145	36%

*Data Courtesy of Levi Rose, Plateau
Action Network

Conclusions

- We can exploit natural biological low-pH oxidation processes as a pre-treatment strategy
- Fe removal enhances the ability of alkalinity generating processes
- Design parameters, such as hydraulic residence time and surface area, can be **optimized**
- Variations in biotic diversity and seasonal variations are poorly understood
- Bioremediation projects are currently being explored as long-term passive treatment options





Thanks for your attention